Non Sibi High School

Andover's Chem 550/580: Advanced Chemistry Chapter 8, Review Quiz 1

1

Rank each group of atoms and ions from smallest to largest radius:

$$\begin{aligned} &\text{a. N, O, P}^{3-}, P \\ &\text{b. Ca, K, Mg}^{2+}, \text{Mg} \\ &\text{c. Br}^{-}, \text{Rb}^{+}, \text{Se}^{2-}, \text{Sr}^{2+} \end{aligned}$$

2

Rank Ar, Ba, Cl, and Cs from smallest to largest atomic radius and lowest to highest first ionization energy.

3

Explain the huge increase in ionization energy between ${\rm I}_4$ and ${\rm I}_5$ for carbon.



 $\frac{\text{Creative Commons Attribution-NonCommercial-NoDerivs 3.0 Unported License}}{\text{Contact: kcardozo@andover.edu}}$